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Organic chemistry dictionary pdf

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Independent, trusted online education for over 23 years! Copyright ©2021 GetEducated.com; Approved College, LLC All Rights Reserved by Christene Barberich such as many a good success story, Vavra Pade's designer jewels began in his garage. I worked in there designer L.A.ki on say. Finally, I had a hard time finding a place for a buyer, let alone a computer. This needs to expand Lead Pade and his assistant, Parish, into a bigger space in downtown Los Angeles. Loft-like and colon, among Little Tokyo and Chinatown and other rooms located in the City Fashion District, it works currently sounding more like a New York environment than L.A. but spending time with the designer's unforgettable creation, and one realizing its approach to jewelry makes is not typical either. I was drawing jewelry on a subconscious level allÁfá €ŠÁ my life,' says Pade. Take starch or stone and string them together. After a brief studying interior design, Pade went to the Gemological Institute where he learned how to perform class diamonds, distinguished stones, cast, and soder. It was also where he learned the practice of bearing jewelry. It's a lost art, indeed,' he said. It's very precise; you work down in the millimeter on paper, so you really learn how to guage size and shape, how to set things up, which watches will work best. It's very traditional to teach you, but then you can take it in your own direction. His own direction was eventually brought back to his early foundation for organic materials. However, today, the designer favors prehistoric shark teeth and 140 million coral on found seashells. The materials I use really inspired the collection,' he said. The sharks' teeth are from the Pliocene era and the Alaskan coral is from the Jurassic period. I like using the things that are really unique andÁfá €ŠÁ this has existed long ago. It reminds us of mortality. It's of course during Pade's and other designs that these ancient materials are giving new life. The crown's teeth with a gold crown and a gold crown and diamond. a delicate piece of rare hanging the designer's rare railroads from a spare gold chain. I love the juxtaposition of material,' he said. I love the contradiction. Treading further in its own direction, the designer creates his own special alloy, a mixture of gold and yellow gold, which yielded a beautiful peach-hued gold 18-karat gold mediumÁfá €ŠÁ to work out. Using this custom metal, the designer has crafted one of her favorite designs (which she's wearing every day), in entirely hand-four ring hummingbird. It's both before all editing and sensual, says Pade. But there's still a feminine foot. The ring of hummingbird is part of the fine jewelry designer's collection. However, it also draws a bridge line called PAdE by Pade Vavra, who has a more playing approach and is seasoned defined. The fine jewelry collection is more consistent aesthetically, he said. I don't want to manipulate things too much, says Pade. I want a piece to look like it came out on earth. Pade Vavra available in TG170 It can favor material that is over 100 million years old, but los Angeles-based Vavra Pade's approach to make jewelry do is nothing but ancient history. This alphabetical chemistry dictionary offers definitions and examples of relevant chemistry and chemical engineering terms. For each term, they provide a short definition. Each link leads to a more complete discussion of the word. Additional definitions are also available. Alkalinity is a measure of how basic a substance is. JazzIRT/Getty Images absolute alcohol – common name for high purity etanol or alcohol ethyl. absolute error – expression of the precarious or precision of a measure. absolute temperature – temperature measured Kelvin Scale. absolute skeptical - skeptical of a scientific measurement, given in the same units as the measurement. absolute zero - the lowest possible state in which problems may exist, 0 K or -273.15 °C. absorb - measures the amount of absorbed light by a sample. absorption – process by which atoms, ions, or molecules enter an essential phase. spectroscopy absorption – technique used to determine the concentration and structure of a sample based on the wave of liquids to absorb. spectrum absorption – graph of the number of absorption as a function of length. absorptivity - absorption cross section of extension coefficient, which is absorbed into a solution for each unit length and concentration. accuracy - accuracy of a measure of a true or accepted value. acid - a chemical species that accepts electrons or donor or hydrogen protons. anhydride acid – a manetal oxide that reacts with water to form an acid solution. Ion-based acid indicators – a weak acid or weak base that changes color when concentration of hydrogen or hydroxide ion changes to an aqueous solution. Acid-based tract – a procedure to obtain concentration of an acid or base does not react to a known concentration and the unknown until the equivalence point reaches. Constant acid – Case – a quantitative measurement of how strong an acid is. acid solution – an aqueous solution with a pH less than 7.0. Consider - Usually considered 90 elements (thorium) of 103 (lawrencium). Otherwise, actinides are defined according to common properties. aktiniom - the name for the element with atomic number 89 and is represented by the symbol Ac. He is a member of the actinide group. activated complex

– an intermediate state of the maximum energy point on the reaction path performed as reactants were converted to produce in a chemical reaction. energy activation – Ea – The minimum amount of energy needed for a chemical reaction occurs. active transportation - the movement of molecules or ions from a region of lower concentration to higher concentrations; require serial energy activity – metal list is sorted in order to reduce activity, used to predict which metal moves others into aqueous solutions. Actual yield – the amount of experimentally generated is found in a chemical reaction. health effects effects – the effect caused by initial exposure to a chemical product. acyl group - a functional group with the RCO formula - where R is bound to carbon via a single bond. adsorption - the memberships of a chemical species on one surface adulatory - a chemical product that acts as a contaminate in the context of purity another substance. another - an average believe brings light waves to the 18th and 19th centuries. when - the mixture of gas that makes up Earth's atmosphere, including of nitrogen, and oxygen, water vapor, argon, and carbon gas. alchim - Several definitions of alchemists exist. Originally, alchim has been an ancient tradition of sacramental chemistry used to discover the spiritual and temporal nature of reality, its structure, law, and function. alcohol - a substance with a group -OH attached to a hydrocarbon. Amino acid - amino acid containing an aliphatic side dog. aliphatic compounds – an organic compound containing carbon and hydrogen together in straight ranges, branch ranges, or non-aromatic rings. aliphatic hydrocarbon - a carbon and hydrogen together in straight ranges, branch ranges, or non-aromatic rings. metal alkali - any element found in IA group (first column) of the periodic table. alcohol – an aqunt solution with a pH greater than 7. alkalinity – a quantitative measurement of a solution's ability to neutralize an acid. a hydrocarbon that has a double carbon-carbon bond. alkanyl group - the hydrocarbon group forms when a hydrogen atom is removed from an alcane group. alkoxide - a functional organic group formed when a hydrogen atom is removed from the hydroxyl group of an alcoholic when it reacts with a metal. group of alkox - functional group that has a bond alkyl group of oxygen. allotrop - a form of an elementary substance. alignment - substance made by melt together two or more elements, at least the one that belongs to a metal. alpha career - spontaneous radioactive displease that produces an alpha particle or nuclear elium. alpha radiation - the oninion release radiation of radioactive disposition of an alpha particle. aluminum or aluminum - the name for the element with atomic number 13 and is represented by the symbol Al. is a member of the metal group. fillings – any of alloy in mercury with one or more metal. radioactive metal and symbol element cores with atomic number 95. amide - functional group that has a carbon group linked to an atom of liquids. amino - composed of which one or more hydrogen in harmony is replaced by a functional organic group. Amino acid - an organic acid containing a carbohydic (-COOH) and amino (-NH2) functional groups together with a side chain. amorphous - theme that describes a solid with no crystalline structure. anfrpotic - species that can be accepted and provide a proton or hydrogen. substance that can act as either an acid or a base. Photo oxide – oxide that can act as either an acid or a basis in a reaction to produce one with water. Zen - the atomic mass unit or the 1/12th mass of an atom with no carbon-12 mass. chemistry analytical - disciplinary chemistry that chemical composition studies of materials and tools used to examine them. angstrom - unit of length equal to 10-10 meters. amount of angular momentum the number of proportions associated with the angular momentum of an electron. anhydrous - describing a substance with no water or others is as concentrated as it can be found. a - an ion with a negative electric charge. anodies - electrons where oxidation occurs; positive charging annihilate orbital antibodies - molecular orbital and an electron outside the region between the two nuclei. anti-Markovnikov addition - an addition reaction between an electrophilous HX compound and either an alkene or alkyne in that hydrogen in the carbon bone and the smallest number of hydrogen atoms and X links to the other carbon. Antimony - Antimony is the name for the element with atomic number 36 and is represented by the KR symbol. It's a member of the metalloid group. anti-periplanar - periplanar compliance where the dihedral room between atoms is between 150°C and 180°. aquent - describes a water-residual system. Aqueous solution – a solution in which water is the solvent. aqua regia - mixture of hydrochloric and nitric acid, capable of gold dissolving, platinum, and palladium. argon - Argon is the name for the element with atomic number 18 and is represented by ar's symbol. This is a member of the noble gas group. aromatic compound – an organic molecule that has a benzene ring. Arhenius acid - dissociated species of water are formed protons or hydrogen. Arhenius base – species that increase the amount of hydroxide an hour to add to water. arbitrary - metalloid and symbols inserted As with atomic number 33. aryl - a functional group from a simple aromatic ring when a single hydrogen is removed from the ring. astatine - Astatine is the name for the element with atomic number 85 and is represented by the symbol In. He is a member of the halogen group. atom – the unit is defined in a component, which cannot be subdivided using chemical meaning. Atomic mass – average mass of atoms of a component. Atomic mass unit (amu) - the 1/12th mass of a carbon-12 mass atom, used to represent atomic and molecular mass. atomic number – the number of protons at the core of an atom an element. atomic radius – value used describes the size of an atom, usually half the distance between two atoms just touching one another. solid atomic – solid in which the atoms are bondage to other atoms of the same type. atomic volume - volume occupied by one mollet of one component at room temperature. Atomic weight – average mass of atoms of a component. atmosphere - fuel environment, such as fuel enthusiasts a planet designed in place by gravity. ATP – ATP is the acronym for triphosphate adenosine molecules. Principle Aufbau – the idea that electrons are added to orbital as protons are added to an atom. austenite - figure-centered crystalline cube form of iron. Avogadro's law - relationship that states are equal traders all gases have the same amount of molecules at the same pressure and temperature. Avogadro's number – the number of particles in one molten of a substance; 6.0221x 1023 azeotrope - a solution that maintains its chemical composition when distilled. azimuthal proportion number - the quantity number associated with angular momentum of an electron, determines the orbital form of its. Boil occurs when the vapor pressure of a liquid pressure exceeds atmospheric pressure. David Murray and Jules Selmes/Getty Images background radiation – radiation from external sources, typically from cosmic radiation and xysop dismay. back attraction – the attraction that the concentration analyzed determines by reacting to a known quantity of reactive excess. Balanced equation – the chemical equation in which the number and type of atom and the electrical load are the same both the reactive and product location of the equation. Balmer series – the hydrogen spectrum part for our electronic transition = 2 and n>2. There are four lines in the visible spectrum. barium - alcohol was metal and low element symbol with 56 atomic numbers. barometer - instrument used to measure atmospheric pressure. base – chemical species that either accept protons or others donate to electrons or hydroxide ion. anhydride base (basic anhydride) – a metal oxide formed from reaction between water and a basic solution. Metal base – any metal besides a precious or noble metal used for jewelry or industries. basic - alcohol or with a pH > 7. basic solution – aqueous solutions with more hydroxide ions than hydrogen hydrogen; solution with pH > 7. Beer's law (Beer-Lambert Law) – state law concentration a solution is directly proportional to absorb its light. berkelium – radioactive metal and symbol element BK with atomic number 97. berillium - alcohol was metal and eBay element symbol with 4 atomic numbers. discouraged beta – the type of radio-active that results in spontaneous emissions of a beta particle. particles - an electron or possit emitron during beta decomposition. radiation beta - radiation slug from decomposed beta in the form of a drastic electron or positron. binary acid – a binary acid compound in which one element is hydrogen and the other element is another manetal. binary compound – a compound made up of two components (e.g., HF). energy required - energy needed to remove an electron from an atom or to separate a proton or neutron to the atomic core. Biochemistry - Biochemistry is the chemistry of the things that live. bismuth - Bismuth is the name for the element with atomic number 83 and is represented by the Bi symbol. It's a member of the metal group. bitumen - natural mixture of aromatic hydrocarbons (PAHs). dark light - a lamp that emits the ulcerete radiation or the invisible radiation emitted by it. Cotlymer - coupon formed by repeated monomer subunits. bohrium – transition metal and symbol element Bh with atomic number 107. Boil - transition phase from the liquid to gas. boiling point - temperature at that pressure vapor a liquid is equal to the external gas pressure. boiled elevation point - the increase of a liquid boiling point caused by adding another compound to it. Body angle - the angle that was formed between two adjacent chemical bonds in the same atom. energy-dissociation - energy is required for homolytically breaking a chemical bond. energy components – the amount of energy needed to break one molecule in component atoms. Proper anthalie – wholesome changes resulting when a single molester is linked to a broken species of 298 K. bond length – the distance of poised between nuclear atoms or the core group that shares a chemical bond. good order - a measure of the number of electrons involved in chemical links between two atoms of a molecule; usually equal to the number of links between the atoms. boron - Boron is the name for the element with 5 atomic number and is represented by the symbol B. Is a member of the semimetall group. Boyle's law - the ideal law that establishes the volume of a fuel is inversely proportional to its absolute pressure, supposed to be constant temperature. alkane dog branch - an alkane with alkyl group sets in the central carbon chain. The molecules are tuned, but the entire C-C bond is the only link. Copper - Copper is defined as a copper alloy and zinc. Bromine - Bromine is the name for the element with atomic number 35 and is represented by the Br symbol. He is a member of the halogen group. Bronsted-Lowry acid – hydrogen yield species. Bronsted-Lowry base – species that accept hydrogen hydrogen in a reaction. Copper - Bronze is a copper alloy, usually containing tin as its main addition. defense - either a weak acid with its salt or others a weak basis with its salt that forms an aquelaze-resisting pH solution. The Celsius scale is a common temperature scale of chemistry. Indeed/Getty Images cadmium – Cadmium is the name for the element with atomic number 48 and is represented by the Cd symbol. It is a member of the transition metal group. Caffeine - Caffeine is a naturally found chemical substance in tea and coffee and added of course. Calcium - Calcium is the name for the element with 20 atomic numbers and is represented by Ca's symbol. He is a member of the metal group Metal Tea. calories - units of thermal energy; The amount of energy is required to increase the temperature of 1 gram to 1 degree C or K in standard pressure. calories – instruments designed to measure heat flow in a chemical reaction or physical change. Capillary action - spontaneous drift of liquid into a narrow tube or pore material. carbon - Carbon is the name for the element with atomic number 6 and is represented by the symbol C. He is a member of the manetal group. carbon - an ion that includes a carbon linked to three oxygen atoms (CO32-) or a compound containing this ion. carbonyl - functional group which includes a carbon dual atom the oxygen bond, C = O. carboxyl group - functional group which includes a dual carbon link of oxygen and salt linked to a hydroxyl (-COOH). catalyst – substance that increases the chemical reaction rate by reducing its activation energy. katant - binding to a component itself via link to covalent, forming a dog or cathode ring - electrode where trimming occurs; usually the negative electroder. cathode rails tube - a vacuum tube with a source of electrons, a fluorescent screen, and the means to accelerate and distort the electron beam. cation - ions with a positive electrical load. Celsius temperature scale – temperature scale where 0°C and 100°C are defined as the freezing and boiling points of water, respectively. cerum – rare earth metal with Symbol element symbol and atomic number 58. cesium - Cesium is the name for the element with 55 atomic numbers and is represented by the Cs symbol. It's an alkali metal group member. wealth number (CN) – values describing the quality of diesel fuel combinations, based on the delay between injection and union. chain reaction – put into the chemical reaction that produces react to another reaction. charge – an electrical charge, a preserved owner of subatomic particles to determine their electromagnetic interaction. The Charles Act – the ideal gas law that establishes the volume of an ideal gas is directly proportional to absolute temperatures, expecting constant pressure. chelate - organic compound formed by linked with a polident ligand in a central metal atom, or the act of forming such a compound. chemical – any problems or substances containing mass. chemical changes – process by which one or more substances are changed to form new substances. chemical energy – the energy contained in the internal structure of an atom or molecule. chemical equations – description of a chemical reaction, including the reactants, products, and the reaction direction. chemical poised - state in a chemical reaction where the reaction reagents and products remain stable over time. chemical formula – expression that states the number and types of atoms in a molecule. Kinetic chemical – the study of chemical processing and reaction rates. Chemical properties - features that can be observed when problems undergo a chemical change. chemical reactions – a chemical change in which reactors form one or more new products. chemical symbol - one- or two-letter representation of a chemical product (for example, H, Al). chemical - lights emitted as a result of a chemical reaction chemistry - study of problems with energy and the interaction between the Cherenkov radiation - the electromagracitic radiation chemistry is emitted when a motion charges particles in a faster dielectric medium than the speed of light in the medium. chiral center - the atom of a molecule linked to four chemical species, allowing optical isomerism. chirality - Chirality or chiral describes an image that is not impossible, such as left and right hand. Usually in term chemistry is used to describe a pair of molecules that have the same formulas, but form a pair of structures. Chlorine - halogen and atomic number 17 with compound symbol Cl. chlorineluorokarbon - A chlorineluorokarbon or CFC is a compound containing atoms of chlorine, fluorine, and carbon. chromatography – technical groups are used to separate mixture elements by passing the mixture into a stationary phase. chromium - Chromium is the name for the element with atomic number 24 and is represented by the cr symbol. A member of the transition metal group. system closed - thermal system in which mass is maintained in the system, but energy can be freely entered or out. koagulation - the college or college of particles, usually in a colloid. cobalt – metal transition that is the atomic number 27 and symbol element Co. coenzyme - substance that works with an enzyme to help its function or initiate its actions. cohere – measures how thoroughly molecular sticks to each other or group together. Collages - an important family of proteins found in humans and other animals, found in skins, cartilage, blood vessels, and tendons. Coligative properties - properties of a solution depend on the number of particles in a solvent volume. colloid – a homogeneous mixture in which particles spread by solving out. law combining gas - laws that establish the ratio of the product of pressure and volume, divided by the absolute temperature, are a constant value. Reaction combination – reactions in which two reactants combined to form a single product. combination - chemical reactions between a gas and oxygenizer that yield energy (usually heat and light). common effects - suppressing the effect an electrolyte has on the ion of another electrolyte that shares a common ion. compound - chemical species formed when two or more atoms form a chemical bond. ion complex - the ion that is a metal central linked to one or more ions or molecules. focus – there is a huge ratio of solvent salts. concentration – an expression of the amount of a substance within a defined volume. condenser - state of change of subjects from vapor phase to liquid phase. Condensation reaction – the chemical reaction in which one of the products is water or harmony, also known as a dehydration condensed formula – the chemical formula in which atom symbols are listed in the order are displayed in the molecular structure, with limited code firing. conductor - materials that allow the flow of energy (e.g., electric vigorous, thermal vigorous). compliant - an isomer that is different from another isomer by rotation around one bond. congener - member of the same group of elements in the periodic chart (e.g., iodine and chlorine). conijge – multiple chemistry definitions, referring to Bronsted acid and base, a compound formed by combining other compounds, or the overlap of p-orbital via a sigma bond. Southern Conjugate – HX, a difference compound from an X base by a proton. conjugate base - the species contain a proton to a south-base reaction. energy conservation - laws that states can change shape but cannot be created or destroyed. conservation of mass - state laws, in a closed system, problems can change shape but not be created or destroyed. variable controller – variables that a scientist keeps constants in an experiment; control or constant variable conversion factor - numerical reports that convert one measure from one unit to another. bond coordinator - covalent corporations between two atoms in which one atom supplies both electrons for the bond. composite coordination - compounds containing one or more link coordinators. coordination number - the number of atoms linked to a central atom. Copernium – radioactive components with Cn symbol and atomic number 112. Copper - Copper is the name for the element with 29 atomic numbers and is represented by the symbol Cu. Is a member of the transition metal group. corrosion - irreversible damage to a material or tissue due to a chemical reaction. Corrosive – has the power to cause compelling chemical damage on contact. The Coulumb law - the law that establishes the force between two charges is proportional to the amount of both charges and inverse proportional to the square in the distance between them. links covalently – chemical links between atoms or ions in which the electrons fear are more or less even shared between them. compound consist of covalents - molecules containing covalent chemicals linked. correlation - half the diameter of the part of an atom involved in a correlation bond. creation - form a scalloped form of exposure to a hypertonic solution. critical point – critical state; point where two phases that matter become indistinguishable from one another. cryogenics – study of problems in very low crystal temperatures – problems in which atoms, ions, or molecules are charged in an order, repeating three dimensional patterns. Crystal field divide – the difference in energy between the gubitals of ligands. Crystallization - solidification of problems in the very order form of a crystal. Kiriyum - radioactive metal and components Cm with atomic number 96. Electricity – rates of flow of electricity. Dry ice is the name for solid carbon dioxide. Jasmin Awad / JeEm / Getty Images Dalton - The relationship declaring total pressure of a gas mixture is equal to the sum of partial gas pressures. darmstadtium - Darmstadtium is the name for the element with atomic number 110 and is represented by Ds symbol. Darmstadtium was formerly known as unnamed with Uun symbols. It is a member of the transition metal group. bond - the corporation between atoms in which one atom provides both electrons for the bond. daughter isotopes - produced formed after a radiosiotop (the parent) underwent radioactive decomposition. two Broglie equations – equations that describe the vague properties of problems, declared as constant equal waves Planck divided by the product of mass and speed. Decanting – method of separating mixture by removing layers of liquid from a precipitation. decomposition reaction – chemical reaction in which one reactive reality of or more products. deflagation – the combination type in which the proliferation of flames is less than 100 cents/s and pressure is less than low 0.5%. Dehydration reaction – chemical reaction between two compounds in which one of the products is water. deliques – process by which a solib substance picked up water steam from the atmosphere to form a solution. delocalize electrons - any electron in an ion, atom, or molecule that is not associated with a particular atom or a single corporal. mass - mass per unit volume. dependent variable – variables were measured (tested) in kg-m/s2). deciding – chemical agents that gather water, often used for drying. desublimation - phase changes from solid steam. detergent - cleaning agent with general structure R-SO4-, Na+, where R is a long alkyl group of alkyl. deamagnetic - not attracted to a magnetic field, generally because the material doesn't have electrons that aren't just. Broadcast - movement of a liquid from a region of higher concentration to lower concentration. dilute – solutions with a small amount of salutation relative to the amount of solvent. dipolate - an electric or magnetic charge separation. Dipolate moment – measuring the separation of two charged electrical opposites. Diprotic acid – acids that can provide two hydrogen atoms or protons per molecule in an aqueous solution. direct proportion – relationship between two variables such that the reports are a constant value. disaccharide - carbohydrates formed when from bail, remove a molecule of water from the structures. reaction considerations – chemical reactions in which the class or a session of a reactant is replaced by one from another reactant. dispropotion - chemical reaction (usually rdx) where a molecular form of two or more products distillate. Reaction to dissociation – the chemical reaction in which a reaction breaks two or more parts. dissolving – a dialyt passing of solution, usually a solid going through the liquid phase. distillation - steam formed by a distillation, which can be condensed into a liquid for collection. distillation – technique of heating a liquid to form a steam, which is cooled separate components into liquids based on volatility or boil. divalent cation – positive ion loading with a value of 2. ADN – outxyribonucleic access, an organic molecule that codes for protein. dual bond - the chemical bond in which two pairs of electrons are shared between two atoms. reaction dual replacement – chemical reaction in which two reactors exchange anions/cations formed two new products using the same A. Dry ice – the solid form of carbon dubernium dumber - metal transition and symbol component DB with atomic number 105. ductile - can stretch to a thread without breaking. dynamic poised – chemical poised between forward and reverse reactions at which the rates of reaction are equal to each other. dysprosium – rare earth metal with Dy element symbol and atomic number 66. Electrons are particular and negative charges that abide by the atomic nuclear warheads. Ian Cumming/Getty Images Efficient Nuclear Load – net charged an electron experience of an atom with multiple electrons. efficiency – drooling or turmoil when fuel evolved by a liquid or solid. reflorescence – process by which a hydrat losses water in hydration. Einsteinium - Movement of gas into a bridge or capillary into a vacuum or another gas. einsteinium - Einsteinium is the name for the element with 99 atomic number and is represented by the Symbol Of East. He is a member of the actinide group. elastic - physical properties of problems that describe the ability to return to original form after deformation. electrical conductivity - measurement of the capacity of a substance brings an electric current. electrical resistance – measures how much a resistance material carries an electric current. electrochemical cells - devices that generate a potential difference between electrodes via chemical reactions. Electrochemist - Scientific studies of reactions and species that were formed in the kides between an electrolyte and a conductor, where electron transfer occurs. electromotive strength - emf – the electrical potential generated by either an electrocheckic cell or altered magnetic fields. electrode - the anody or cathode into an electric cell. electrolysis - actual direct passage of a solution, producing a chemical change in the electrodes. electrolytes - a substance that forms ions in aqueous solution. electrolytic cells – the type of electrochemical cells in which the flow of electrical energy from an external source allows a redox reaction. electromagrematic radiation - light self spread energy that contains electric and magnetic field elements. electrons – stable negatively packed subatomic particles. electrons aference – measured in the ability of an atom to accept an electron. Electronic capture (EC) – the radioactive form of the capability that atomic nucleus absorb a K or L electron cult, converting a proton into a clean. cloud electrons – regions in negative charges that enclose the nuclear atomic entrance that has a high probability of having electrons. electronic configuration - description of the population of the electronic energy sublevels in an atom. electron density – representation of the probability of getting an electron in a specific region around an atom or molecule. domain electrons – the number of pairs of electrons lone or places linked around an atom or molecule. electronegativity - the owner of an atom that reflects its ability to attract electrons to a chemical bond. repulsion electronic fear – electron-fearing principles that enthusiasm a oriented atom oriented themselves as far as possible; used to predict geometry. Electronic-sea models - models of metallic bonding in which cations are described as fixed points in a mobile sea of electrons. turn electrons - the owner of an electron related to its turn on an axle, described by a proportion number as either +1/2 or -1/2. electrophil - atoms or molecules that accept an electron pair to form a kovalan bond. electroplating – process of adding a metal coat to a material using a reduction reaction. electrostatic forces - strength between particles due to electrostatic charges. electric - a natural asian in gold and silver. element - a substance that cannot be subdivised using chemical medium; identified by the number of protons in its atoms. Elementary reaction – the chemical reaction that reacts form generated at a single step without a transition state. symbol elements - the two- or two-letter abbreviation of a chemical element (e.g., H, Cl). emissions - produced in a combination reaction, aside from heat and light (e.g., carbon dioxide). Spectrum emissions – range of wavelenocks emitted by an atom encouraged by electricity or heat. embricke formula – formula showing the ratio of elements in a compound but not necessarily the actual numbers of a molecule. stabilized agents that prevent moistive liquids from separating. Emulsions - cologne formed in two or more moist liquids where a liquid has a dispersion of the other liquid(s). inanium – a member of a pair of optical isomers. orthochemic - absorbed process energy in its environment. inediol – an inol alcohol alcohol with a hydroxyl cluster attached to both carbon atoms in the C=C bond. energy - the ability to perform tasks (e.g., kinetic energy, light). enthusiasm - the thermodynamic property of a system that is the sum of the internal energy and the product of pressure and volume. change enthusiasm – the energy change in a system of constant pressure. analysis of atomization – the number of healthy changes when chemical links are broken into a compound to form individual atoms. analysis of reaction – difference between total analysis of products and the total analysis of reactions in a chemical reaction. enterprises – measures of the disorder within a system. Enzyme - An enzim is a protein that functions as a catalyst for a chemical reaction. constant equilibrium – reports of concentrations of poised to produce raised from the power of the stoicometric coefficient at the concentration poised in reactants are raised to the power of their stoicometric coefficient. point of equivalence – point to a titration where the titran completely neutralizes the analysis. Erbium - Erbium is the atomic number 68 component on the periodic chart. Essential acid - amino acid is needed in the diet because an organism cannot synthesize it. ester - RCO2R', where R' is the hydrocarbon part of the carboic acid and R' is the alcohol. ether - organic compounds containing two aryl groups or alkyl tied to an oxygen, R-O-R'. Europium - Europium is the name for the element with atomic number 63 and is represented by Eu symbols. He is a member of the lanthanide group. eutectic – solid homogeneous mixture of at least two types of atoms or molecules that form a superlatice (usually a mixture of alloys). Evaporation – process is characterized by a spontaneous transition of molecules from phases to liquids in steam phases. excess reactant - reactive left over in a reaction because it is present in a large number than needs to interact with the limit limit. excited state - atoms, ion, molecules, or sibatomic particles at an energy level above its soil state. spiritual - releasing energy into its environment. spiritual - releasing energy to the environment in the form of heat; a type of hexagonal reaction process – a chemical reaction that releases heat. Extensive properties - properties of problems depending on the amount of issues that are present (e.g., volume). The blaze test is an analytical technique used to help identify metal ions. (C) Philip Evans/Getty Images orbital - orbital electronics with it = 3 for the angular proportion number, family - a group of elements that share similar properties. Constant faday – a constant physics equal to the electric load of one molester of electrons, 96485.33 C/mol. fat – the sadness of glycerol and fatty acids containing many solutions to organic solvents, but insolub in water. Gracie acid - a carbon acid with a long hydrocarbon side chain. feedstock – any material that does not have processes used as a reserve for a manufacturing process. Fermium – Fermium is the name for the element with 100 atomic numbers and is represented by the symbol fm. He is a member of the actinide group. first law of thermodynamics – laws that establish a system's total energy and its environment are a constant value; conservation law of energy, fiery point – the lowest temperature a steam will begin with sustained combination. fission - the divide of an atomic nuclear, which results in two or more nuclear light with a release of energy. Test flames – an analytical technique used to identify identified based on the emissions spectral of a flame. flammable – easily ignored or capable of sustaining combination. liquid - a substance that flows under stress shear applied, including liquids, gas, and plasma. Fluorescence - imination dropped when a radiation absorbed electromagnitive atoms and emitted a photon when the electron falls into a lower energy state. foam - a substance with gas bubble trapped in a liquid or solid. strength - a push or pull on a mosquito, with both greatness and direction (vector). formal charge – the difference between the amount of electron valence in an atom and the amount of electrons associated with the atom (e.g., in a chemical bond). Reaction formations – reactions in which one mole is in a formed product. mass formulas or weight formulas – the sum of the atomic weights in the atoms of a compound empirically formula. fractional distillation - processes that separate components of a mixture according to their boiling point. Francium - metal alkali and symbol fr elements with 87 atomic numbers. free energy – the amount of internal energy in an available system to do work. free radicals - an atom or molecule and an electron that does not live. Freeze - process in which a liquid changes to a solid. freeze dots – temperature at which a liquid transition to a solid (not always the same as melting points). Freeze Point Depression - lowers the freeze point of a liquid by adding another compound to it. frequency - the number of times a point over a wave than a reference point at one second. functional or functional moiety group – group atoms of a molecule that is responsible for characteristic reactions and properties. Fusion - combined nuclear atomic light to form a heavier nuclear, accompanied by the release of energy. Test tubes are a common type of glassware chemistry. Kultura Science/GiPhotoStock/Getty Images gadolinium – rare earth metal with Gd element symbol and atomic number 64. gallium - metal and symbol ga elements with 31 atomic numbers. galvanic cell - electrochemical cells where reaction between dissimilar worsheteorator occurs in a single bridge and radiation - high energy to insufficient photons, backgrounds from the atomic nucleus.gas – the state of materials characterized by neither a defined form nor defined volume.gas constant (R) - the constant of the Ideal Act; R= 8.3145 J/mol.K. Gay-Lussac - the ideal form that establishes the pressure of an ideal gas is directly proportional to its absolute (Kelvin) temperature when volume is designed constant.gel - a type of soil where the solid particles are made of a mesh to form a rigid or semi-rigid mixture.geometric isomer - molecules which have the same quantity and types of atoms as each other, but with different geometric setup. Also called cis-trans or configurational isomerism.germanium - metalloid and symbol war elements with atomic number 32.Gibs free energy – a potential measurement of reversible or maximum work done by a system of constant pressure and temperature. glass - an amorphous solid. glycosidic bond – a covalent between a carbon and a functional group or another molecule.dog – yellow-colored transition metal and symbol of Zen Number element and atomic number 79.Graham's – relationship that declares the efficiency rate of a gas is inversely proportional to root in squares at its time they mass or alcohol density -purifying ethyl alcohol forms made from distilled fermented grain.grams - equal mass units of the mass of a cubic centimeter in water at 4°C.grams molecular mass – the mass of one gram of a molecular substance. gravimetric analysis – a set of quantitative analytical techniques based on measures of mass mass chemistry samples – branches of chemistry concerned and reduce the environmental effect of chemical products, including the development of new materials and processes.ground state – the lowest energy state of an atom, ion, molecule, or sibatomic particles. group - a vertical column on the periodic table including elements that share periodic properties. Heat refers to thermal energy. Tim Robberts/Getty Images Haber process – method of making harmony or repairing nitrogen by reacting nitrogen and hydrogen fuel entertaining – metal transitions and symbols Hf elements with 72 atomic numbers. half cell - half an electrolytic or voltaic cell, serves as the site of either oxidation or reduction. half-life (1/2) – time required to convert half of the reactant to a product or time needed for half of a radioactive isotop to decompose his daughters' isotopes. Hide ion – a single atom, which has a load of -1 (e.g., Cl-) flows - an element of VIIA group in the periodic table (e.g., Br, Cl). halogen hydrocarbon – a hydrocarbon with one or more halogen atoms. hard water - water with high amounts of calcium and/or magnesium grade. acidium - transition metal that is 108 atomic numbers and symbols of components Hs. heat - energy that halogen sample subjects due to a temperature difference. heat capacity – the amount of heat needs to increase the temperature of a sample by a specified amount. Heat of training (ΔHf) – amount of heat absorbed or released during training of a pure substance in its elements of constant pressure. Heat Fusion (ΔHfus) – healthy changes (heat) for the conversion of a gram or mole a solid to a liquid at constant temperature and pressure. heavy metal - a dense metal that is toxic to low concentration. Heisenberg's skeptical principle – the principles that states it is impossible to determine both the position and momentum of a particle at once and perfect accuracy. Helium - Helium is the name for the element with atomic number 2 and is represented by the symbol Li. The Henderson-Hasselbalch equation – an approximation that relates the pH or pOH to a solution, the pKa or pKb, and the ratio of concentrations of dehydrated species. Henry's law – the state laws of a gas that will dissolve into the solution is directly proportional to the partial pressure of the above solution gas. heterogeneous – including dissimilar elements. heterogeneous mixture - a mixture that lacks a uniform composition such as that at least two elements are present and identified properties. Heterogeneous reaction – the chemical reaction that reacts to different phases of each other. holmium - rare metal earth with ho element symbol with atomic number 67. homogeneous - uniform to its volume. otherwise - polymer in which each merged unit is identical. orbital hybrid - orbital - formed by the combination of two or more orbital atoms. Hydration reaction – reaction in which a hydrogen and hydroxyl ion are attached to a carbon at a C-C dual bond. hydrocarbon - molecules containing entirely of carbon and hydrogen atoms. Hydrogen - inserted with atomic number 1 and symbol H. hydrogen bond - attracts interaction between a hydrogen link to an electronic atom and a different electronic atom. Hydrogenation – hydrogen reduction reactions that produce hydroden (usually as H2). hydrolysis – decomposition reaction in which a reactant is water. Reverse in a condensation reaction. hydrometer - instruments used to measure the relative density of two liquids. Ironom ion – H3O + the cation. hydrophobic - properties repeatedly water. hydroxyl group - functional group which includes an atom hydrogen that is properly covered in an oxygen atom (-OH). igroscopic - can absorb or adsobe water from the environments. hypertonic – has higher osmotic pressure than another solution. hypothesis - prediction of an event or propose an event's explanation Non-mixed fluids are said to be moist. Samborski/Getty Images ideal gas - Fuel in that molecular has negligible size and kinetic energy depending only on temperature. ideal gas constant – physical constants in the Ideal Gas Act, equal to constants in Boltzmann but with different units. ideal gas – PV = nRT where P is pressure, V is volume, we are the number of moles, R is the ideal gas constant, and T is temperature. moistished - owners of the two substances could combine to form a homogeneous mixture; unable to mix independent variables – the variable that controls or changes from an experience to test its effect on the dependent variable. indicator - substance undergoing a visible change when its condition changes (e.g., a pH indicator). indium - metal and symbols of the element with 49 atomic numbers. Productive effects – the effect a chemical bond has on the orientation of bond adjacent to a molecule. ban – substance that slows or prevents a chemical reaction. inorganic chemistry - study of chemistry in molecules of non-biological origins (no C-H bond). insoluble - can melt in a solvent. Intensive properties - properties of problems that are independent of the number of problems in a sample. forces of intermolecular - sums all strength between neighboring molecules. internal energy energy – total energy (U) in a closed system. intrinsic property – property problems that are independent of the number of questions present. intermediate - substance formed in a middle step between reactant and final products. inverse proportion – relationship between variables so that the products are a constant value. iodine – iodine is the name for the element with atomic number 53 and is represented by the symbol I. He is a member of the halogen group. ion - atoms or molecules with a different amount of protons than electrons and thus an electrical net load. ionic - regarding bringing an electrical net load to the atomic or molecular level. Ionic link – chemical link between atoms caused by electrostatic force between the opposite of charging ions. Ionic compound - compound formed by bonding bonding together due to electrostatic strength (different electronic values). Ionic equations - chemical equations in which electrolytes in aqueous solutions are written as dehydious ions. ionic rays – half the distance between two one just touching each other. energy dissolving - energy is required to completely remove an electron from a gas atom in ion. iridium - Iridium is the name for the element with atomic number 77 and is represented by the symbol Ur. Is a member of the transition metal group. structure with thus the same amount of electron valence. isolated system – thermodynamic system that cannot exchange energy or problems outside of the system. isomer - chemical species with the same quantity and type of atom as another species, but a different arrangement and thus different properties. Process isomerization – protocols in which straight chain hydrocarbons are converted into hydrocarbons chain branches. isotopes - atoms containing the same number of protons, but different numbers of neutrons and thus different atomic weight values. IUPAC – International Union of Pure and applied chemistry, an authority on chemical standards. The newspaper is an energy unit. Paper Vintage Creative/Getty Images Joule – SI energy unit equal to kinetic energy in a 1 kg mass moving to 1 m/s. Crypton is a noble gas. Science Photo Body/Getty Images Kelvin Scale temperature – an absolute temperature scale with 100 degrees between points to freeze and boil in water (although values are provided without degrees per convention) paste - a fiber protein produced by cordate. It can be found in hair, skin, glitter, and quilt. ketone - composed of a functional carbon diet group (C=O) between the two kilo atom groups – prefix meaning a thousand. kilopascal (kPa) - pressure unit that exists by a 10 g mes on a square centimeter. There is 1000 Per 1 kPa. energy kinetics - energy associated with movement. Krypton - 36 elements on the periodic table with symbol Kr. Litmus paper is a specific type of pH paper. Clive Streeter/Getty Images labile complex – a complex ion that quickly reacts poised and ligands in the entourage solution. lanthanides - subsets of transition metal characterized by filling in the 4f sublevel, usually atomic number 58-71. lanthanum - the number of atomic elements 57 and symbols of energy latency elements - changes analyzed in the process by opposing-charging areas of a gas combine to form a solid latils. law - a general rule that explains a scientific observation body. Laws are declared in words, but expressed by mathematical equations. Chemical poised law – an expression of the relationship between the concentration of reactants and the product of a chemical reaction mixture in poised. Law volumes – Relationships that state the fuel volumes of a chemical reaction are present in the ratio of small integers under conditions where all gases are at the same temperature and pressure. Conservation laws of energy - laws that state energy cannot create or destroy, although it can be changed from one form to another. Conservation law in Mass - state laws that matter to a closed system can neither create nor destroy, although it can change shape. Composition Law Constants - Chemistry law that samples the states of a better quality compound have the same elements of the same March. Definite proportion law – laws that state all samples of a compound have the same proportion of elements by mass. Multiple proportion laws - laws that state combined to report in whole numbers gradually form molecules. lawrencium - aggravated with component symbol Lr with atomic number 103. lead – metal and symbol pb elements with 82 atomic numbers. When Chatelier' Principles - the principle that is poised in a chemical system will shift towards stress relieve. Lewis Aside - chemical species that can act as an accepted pair of electrons. Lewis Base - a substance that is an electron donor. Lewis resolute base reactions – chemical reactions that form at least one correlating bond between an electron peaceful donor (Lewis base) and accepting electronic fear (Lewis Aside). Lewis structure – representation of a molecule that uses points to show electrons around atoms and lines to display consituent links. ligand - a chemical species that provides or shares at least one electron via a corporation bond with a central ion or atom. limited reaction – the reaction that determines how much product can cause a chemical reaction. lipid - grade of fat-soluble molecules, also known as oils and liquid liquid thanks – processing of converting a material from a solid phase or gas to the liquid phase. liquid – state of subjects characterized by having a definite volume, but not a definite form. lithium - alkali metal with 3 atomic numbers and 3 component symbols Litmus paper – filter paper used as a pH paper that was treated with a water-soluble found in lichen. London dispersion forces – weak intermolecular forces between atoms or molecules in close proximity to each other, due to electron repulsion. pair of connections - an electron pair of the outward steroid of a non-shared atom or bond with another atom. lithium – rare earth metal with component symbol Lu and atomic number 71. Mass is a measure of the number of subjects in a sample. Larry Washburn/Getty Images macromolecule - molecules with a large amount of atoms, mainly more than 100. The Telung Regulation - rules that describe the filling of orbital electrons in atoms due to the safeguarding of nuclear charges by inner electrons. Magnesium - Magnesium is the name for the element with 12 atomic numbers and is represented by symbol Mg. Magnesium is an alkali earth metal, usually group element – any of the elements of the s and block of the periodic table. mailable - can be shaped or plunged with a hammer, usually applied to metals. Manganese – components with 25 atomic numbers and symbols elements Mn. manometer - device used to measure gas pressure. mass - amount to issue a substance there or owner of problem-resisting acceleration. mass damage – the difference between the masses of an atom and the mass sums of its protons, neutrons, and electrons. number - integer number which is to sum the number of protons and neutrons in the atomic nuclear. mass rate – concentration calculated as mass of a component divided by total mass of mixture or solution; w/ % mass spectroscopy – scanned techniques used to separate and/or identify components of a mixture based on mass and electric charges. questions - anything with mass and busy volume. measure – quantitative or numerical data describing an object or event. chemistry medicine - branches of chemistry concerned with design, synthesis, and study of pharmaceuticals. radioactive transition metal and symbol Element Mt with atomic number 109. dissolved - phase changes in subjects from liquid solids. melt point – temperature in which the solid and liquid phase of the coexist subjects is poised. mendeleevium – agitated with 101 atomic numbers with component symbol Md. meniscus - boundary phase between a liquid in a container and a gas, curve due to surface tension. mercaptan - organic sulfur compound containing an alkyl or aril group and a thiol band. mercapto group - functional group which includes a linked flow of a hydrogen; -SH. mercury - transition metal and symbol Hg element with atomic number Hg. firmware - put in biochemical reactions that store chemical energy and convert it to a form an organism can use. metal - substances with high conductivity and other metallic properties, including trend form cations, often identified by groups on the periodic table. metallic characters - set of chemical properties associated with metal, including the ability to lose electron external valens in form of cations. Metallic compound – chemical compound containing one or more metal atoms. metalloid - inserted with intermediate properties between those of metal and nonmetals (e.g., silicone). meter - either (a) the base unit in length of the system IF or (b) a device used to measure a quantity. Methyl group - functional group that has a carbon link to three hydrogen atoms, -CH3. microliter - volume unit which is a million dollars a liter a cubic millimeter. micron - unit length equal to a million dollars in a meter; a micrometer. Mineral acid – any inorganic acid (e.g., sulfuric acid). miscible - dissolvable or capable mixing to form a solution, typically applied to liquids. mixture - combination of two or more substances such as that each maintains its separate chemical identity (e.g., salt and flour). moderators – material that slows or moderates the speed of neutrons. Mohs Scale – Mohs scale is a relative scale rating the hardness of a mineral. A mineral with a high number Mohs is able to mark a mineral with a lower Mohs number. moiety - group of atoms in a molecule that is responsible for its characteristic chemical behavior. molality - units of concentration which are the moles of salt by kilograms are in solvent. molar - refers to molarities (moles per liter of solution); e.g. a 6 M HCl solution contains 6 moles of hydrochloric acid per liter of solution. Molecular analysis of fusion – energy needed to change one molester of one substance from solid liquid phase to constant pressure and temperature. Motivated analysis of vaporization – energy needs to change a single molecular in liquids in the gas phase of constant pressure and temperature. molarity – unit of concentration which is the number of sealed moles divided by number of liters of solution. Molten mask – the mass of one molle in a substance. Heat abilities molten - energy heat is required to increase the temperature of 1 molester to a 1 substance of Kelvin. motivated volume - the volume of one mollet of a substance. molecules - Chemical mass unit is equal to 6.022 x 1023 molecules, atoms, or other particles. molecular equations – the balanced chemical equation in which ionic compounds are expressed as molecules rather than ions. molecular formula – expression of the number and type atom of a molecule. molecular geometry - description of the form of a molecule and the relative positions of its atom. Molecular mass – sum of atomic mass of atoms in a molecule. molecular orbital - wave function of an electron in a molecule. molecular weight – sum of atomic weight of atoms in a molecule. molecules - chemical species formed by two or more atoms that share chemical links like they form a unit. molelure fraction – unit of concentration which is the number of moles of a component divided by the total number of moles in a solution. molecular report – reports or fractions compare the numbers of moles of any two elements involved in a chemical reaction. mobdenom - transition

metal and word element symbol with atomic number 42. monatomic ions - an ion formed by a single atom. monomer - a molecule that is a subunit block or building of a polymer. Monoprotic acid – acid that provides a single proton or hydrogen atom for each molecule in aqueous solution. luxury mothers – crystal remaining solutions removed from a crystallization solution. MSDS - Acronym for Material Safety Sheet. a written information about safety written on a chemical. bond is formed when two or more pairs of electrons are shared between two atoms. muriatic acid, HCl. Neon light has the noble gas fully. Jill Tindall/Getty Images Naphthenes – cyclic aliphatic hydrocarbons from petroleum and the general formula CnH2n. Natural abundance – average rate of a naturally isotope occurring on Earth. neodymium - rare earth metal with Nd element symbol and atomic number 60. net - noble gas with nerve element symbol with atomic number 10. neptunium – adjudicated with NP component symbol with 94 atomic number. Ionic neutral equation - the chemical equation listed only participating in the reaction. solid network - material that includes an array of repeated concrete atoms. neutral solution – aqent solution with a pH of 7. Neutralization – Chemical reactions between an acid and basis that result in a complete solution. Neutron - particles of the atomic nuclear containing a mass of 1 and load of 0. NEWTON(N) – IF strength unit is equal to the amount of strength needed to accelerate a 1 kg mass 1 m/sec2. Nickel - Nickel is the name for the element with atomic number 28 and is represented by the symbol Ni. Nickel is a member of the Transition Metal Group. niobium - Niobium is the name for the element with atomic number 41 and is represented by symbol Nb. Niobium is also called Columbium and is a transition metal. Nitrogen - Nitrogen is the name for the element with atomic number 7 and is represented by symbol N. Nitrogen is also known as nitrogen and is a member of the mantal group. Nobelium - aggravated with element symbol Name and atomic number 102. Noble gases - components from Group 8 to the periodic chart (e.g., xenon, argon). noble gas – sorthand notation used to record atomic configuration of which is before noble gas setup is replaced by the element symbol in brackets. non-electronic - electronics in an atom that is not involved in a chemical bond with other atoms. cytrolytes - non-dissociated substances in ion of aqueous solutions. mantal - element that does not show metallic properties, typically refers to elements located at the upper right corner of the periodic chart. nonoxidizing acid – an acid that cannot act as an oxidized agent. nonpolar bond – chemical bond with the same load distribution that it does not have positive or negative poles. nonpolar molecules – molecules that have the same load distributions that it does not have positive and negative places. Nonspontaneous reactions – chemical reactions that cannot occur without input from external tasks. nonvolatile - substances that are not easily evaporated in a gas under ordinary conditions. normal boiling point – temperature at which a liquid boiled to 1 at pressure (sea level). normal concentration – either refers to normal concentrations in which the concentration of solites is the same in two samples or refers to weight grams equivalent to a solution of solution (N). normality (N) – concentration measurement equal to weight equivalent grams per liter of solution. normal melt point – temperature in which a solid melt in 1 atm of pressure. nuclear fission – divided into atomic nuclear into two or more nuclear light, accompanied by an energy release. naked radiation – particles and photons emitted during reaction to the atomic core. Nucleation - processing of steam condenser droplets in a liquid, bullet formed in a boiled liquid, or particle crystals. nucleophile - atoms or molecules that provide an electron pair to form a kovalan bond. nucleotide - organic molecules consist of a nuclear base, riboss or deoxygen, and one or more cysphate groups. Nuclear – positively charged center in an atom, made from protons and neutrons. nuclide - an atom or ion characterised by its nuclear. Null hypothesis – proposition that there is no effect in a treatment or no relationship between an independent and dependent variable. nutraceutical – a food or part of food that confers health or medical benefits. Two oxygen atoms to form an oxygen molecule. ADAM HART-DAVIS / SCIENCE PHOTO LIBRARY / Getty Images octane numbers - values indicating resistance to engine fuel the engine hit relative to the hit from isooctane (100) and heptane(0). octet - group 8 valen electrons around an atom. octet rules – principals who atoms in an atomic bond share of their 8 electrons. open system – a system able to freely exchange subjects with energy and its environment. orbital - Mathematical function that describes the vague behavior of an electron. organic chemistry - studied in the chemistry of the compound containing the carbon chemical bond of hydrogen. osmium - Osmium is the name for the element with atomic number 76 and is represented by the os symbol. It is a member of the transition metal group. osmosis - movement of solvent molecules via a semiperable membrane from a dilute solution to a more focused solution, thus diluting it with equal concentration on both sides of the membrane. Oxidation – loss of electrons by an atom, molecule or ion in a chemical reaction. oxidation numbers – the electrical load of a central atom in a compound coordination if all electronic pairs and ligands have been removed. Oxidation state - the difference between the number of electrons in an atom of a compound compared to the number of electrons in a net atom of the element. oxide - an ion of oxygen and an oxidation state equal to 2- (e.g. oxide done). oxidizer - a reactant that removes electrons from another reaction to a redox reaction. oxidized agent - an oxidized; a reactant that removes electrons from another reactant. Oxygen - an anion containing the element's oxygen. oxygen - Oxygen is the name for the element with atomic number 8 and is represented by the symbol O. He is a member of the mantal group. Periodic charts organize elements according to trends in the properties. Digital Art/Getty Images palladium – metal transition with component symbol Pd and atomic number 46. paramagnetism - material properties characterized by being attracted to a magnetic field. atom - radioactive atoms of one or more atom daughters. nuculation of parents - nuclide that careers to a specific nuclide daughter during rotting radioactive. Partial pressure – the pressure of a gas in a mixture of gas would exercise if it handles the volume alone, at the same temperature. parts per million (PPM) – units of concentration which are part solution for every million parts solvent. Pascal (Not) – if pressure unit is equal to force 1 Newton per square meter. Pauli exclusion principle - principles that say no two electrons or other farmland may have identical proportions numbers in the same thing or molecules. percent composition – percent by mass of each component of a compound. yields percent – ratio rates of real yield divided by theoretical yield. perriplanar – describes two atoms or groups of atoms in the plane just as each other with respect to a single bond. period - horizontal range in the periodic table; components with the energy level even higher electronic inexcisities. periodic law - laws that state the property of repeated elements in a practical and systematic way when they arrange by increasing atomic numbers. periodic chart - the inserted chart arrangement by increasing atomic numbers, ordered according to trends of recurring propertes. periodic trend – regular variations of the properties of elements and increasing atomic numbers. periodic – recurring variation of element properties and increasing atomic numbers due to trends in atomic structure. peroxide - a poliotomic spine and O22-molecular formula. petroleum - oil break; natural flammable hydrocarbon mixture found in geologic formation. pH – measurement of concentration of hydrogen, reflecting how acid or basic a substance is. phase – different forms of problems with uniform chemical and physical properties. phase change – changes in the state of the problem of a sample (e.g., steam liquid), phase diagram - chart showing phases of a substance according to temperature and pressure. phenolphthalein - an organic pH indicator, C20H14O4. pH indicator - compound that changes color on a range of pH values. Phlogiston – Filogistn – Filogist is believed to be a substance all problems combined with war when consumed. Phlogist theory was an early chemical theory to explain the process of oxidation. Filogist had no twist, flavor, color or mask. The substance of the deflojst called the calx of the substance. pH meters – instruments that measure pH in a solution based on the voltage between two electrodes in the solution. metaphorically - libels generated when electromagnically energy (usually UV light) kicked an electron from a lower pan in higher state of energy. A photon is released when the electrons fall in a lower state. phosphorous - mantal and symbol inserted P with atomic number 15. electromagromagnetic radiation packages. physical changes – changed the shape of problems but not its chemical composition. Physical properties – the problem feature that can be observed and measured without changing the identity of the sample. pKa – negative base 10 liter logs of constant acid acid; the lower pKa correlates are stronger pKb acid – negative base 10 liters the constant base base; lower correlate pKa with stronger base. 6,626 x 10-34 J·sec. plasma - state with no defined form or volume including ions and electrons. platinum - metal transition with 78 atomic number and symbol element Pt. plutonium - Plutonium is the name for the element with 94 atomic number and is represented by pu’s symbol. He is a member of the actinide group. pniktogen - member of the nitrogen element group. pOH – measurement of the concentration of hydroxide ion in an aqueant solution. polar ties – the covalent types of convalents in which electrons are not equally shared between the atoms. polar molecules - molecules that have polar links such as that sum of the moment to dipover bond is not zero. polonium - atomic number 84 element and skin element symbol. Polyatomic ion - composite ion of two or more atoms. polymer - large molecules made of rings or ranges in repeated monomer subunits. polinuclear aromatic hydrocarbon - hydrocarbon made of aromatic ring collar. Polyprotic acid - acid can provide more than one hydrogen atom or proton per molecule in an aqueous solution. positron – the antimatter counterpart of an electron, which has a charge of +1. potassium – metal alkali and symbol element K with atomic number 19. potential difference – work is required to move an electric charge from one point to another. energy potential – energy due to an object of position. PPB - parts per billion PPM – parts per million praseodymi – rare earth elements with Pr symbols and 59 atomic numbers. precipitation - to form an insolub compound does not react to salt or change a compound’s solubility. Precipitation reaction – chemical reaction between two soluble salts in which a product is an insoluble one. Pressure - strength measurement per unit of the area. main standard - very reactive is pure. primary energy level – primary energy signature of an electron, indicated by the amount of our proportions. principal proportion number - our proportion number describing the size of an electronic orbital. product - substance formed as a result of a chemical reaction. promising – rare Earth element with 61 atomic numbers and pm element symbols. proof - the percentage volume of ethyl alcohol in an alcoholic beverage. properties – the issue feature fixed by its state. protaktinium - magnesium atomic number 91 and symbol inserted By. proton - component of the atomic nuclear with a mass defined at 1 and load of +1. protonation - addition of a proton to an atom, ion, or molecule. PSI – pressure unit; square inch. pure substance – samples of problems with constant composition and different chemical properties. Qualitative analysis determines the composition of a sample. Rafe Swan/Getty Images qualitative analysis – determination of chemical composition in a quantitative sample analysis – determination of the amount or number of components in a sample. quantum – a discreet package of problems or energy. pluable is the number of proportions – values used to describe the energy levels of atoms or molecules. There are four proportions numbers. Radiation refers to any form of uniformed energy. Mads Perch/ Getty Images Radiation - energy emitted in the form of radius, waves, or particles. radio- spontaneous emissions of radiation as particles or photons from a nuclear reaction. radioactive tracer – radioactive components or compounds added to a material to control its progress in a system. Radiom - Radiom is the name for the element with atomic number 88 and is represented by the Rare symbol. He is a member of the alcohol metal group Dia. radon - radioactive gas with Rn element symbol and atomic number 86. Raoult’s law – the relationship that establishes the vapor pressure of a solution depends on the molar fraction of salt added to the solution. reactant – starts material for a chemical reaction. reaction - a chemical change that forms new substances. Reaction quotient – Q – reports of concentration of products in a reaction to the concentrations of reactions. reaction rate – the speed at which chemical reactor forms the products. reactive – compound or mixture has been added to a system to produce a reaction or test if a person occurs. real fuel – fuel that does not behave as an ideal gas because its molecules communicate with each other. indicator redox - compound that changes color to a specific potential difference. reaction redox – put into chemical reactions involving reduction and oxidation redox titration – titration of reducing agents by an oxidized agent or vice versa. reduction – half the reaction in which a chemical species decreases its oxidation number. generally by taking electrons. Refrigerant – compounds that easily absorb heat and disclose it to a higher temperature and pressure. relative density – reports of density of a substance with the density of water. Relative error – skeptical of a measure compared to the size of the measurement. relative to standard deviation – measurement of the data accuracy, calculated by dividing standard deviation by average of data values. relative skeptical – relative errors; uncertainty a measure compared to the size of the measure. residue - questions left after or distillation or an unruly reaction or a recognizable portion of a larger molecule. reason – average in two or more Lewis structures, different from the electron position. reverse osmosis – filtering methods that work by applying pressure on one side to a reaction membrane semiperable reaction – the chemical reaction in which the chemicals act as reactive for the reverse reaction. rhenium - metal transition with 75 atomic number and symbol of components Re. rhodium - metal transition with 45 atomic number and rh element symbol. RNA - riboncule acid, a molecule that codes for sequence amino acid. Roasting – Metallic processing in which a scheduled sulfide ore is heated in when forming a free metal or metal oxide. roentgenium - radioactive components with 111 atomic numbers and symbols inserted Rg. room temperature - comfortable temperature for people, typically around 300 K. RT - abbreviation for room temperature; air-conditioning temperatures are comfortable for people. Rubidium - Rubidium is the name for the element with atomic number 37 and is represented by the Rb symbol. It’s an alkali metal group member. ruthenium - metal transition with 45 atomic numbers and symbols components Ru. petroleum – radioactive transition metal and symbol element RFC with 104 atomic number. Gallium is an example of a semimetal. Science Photo Co/Getty Images Single - Ionic compound formed by reacting an association with a base; sometimes referred only to sodium chlorides, NaCl. single bridge - connection with a weak electrolyte located between the oxidation and half cell reduction of a galvanical cell. samarium – earthy components with 62 atomic numbers and symbols inserted Sm. saponification - reactions between triglycerides and either sodium sodium or hydroxide potassium are formed a fatty acid called soap and slippery. saturated – either a substance in which all atoms are linked by one linked one, a solution with maximum solid melt concentration, or a wet material. saturated fat – lipids that have only one C-C bond. Saturated solution – chemical solution with maximum concentration of melted solutions for that temperature. scandium - Scandium is the name for the element with atomic number 21 and is represented by the sc symbol. It is a member of the transition metal group. science - the systematic study of the nature and behavior of the world using observation and experience scientific law - general rules that explain an observation body in the form of a mathematical or verbal statement and involves a cause of a relationship effect between observations. scientific methods – the system gaining knowledge and solving problems through observation and experimental testing of hypothesis. seaborgium – radioactive transition metal with component symbol Sg and atomic number 106. Second number – read, the quantity number associated with the angular momentum of an atomic electron. selenium - numberetral and symbol elements ls with atomic number 34. semi-metal - inserted with a partially full orbital p, resulting in intermediate property display between metal and mantetal ones. SI – International system, the standard metric system of units. sigma link - conveniently formed by overlap of orbital exterior of adjacent atoms. simplest formula – reports of elements in a compound. only reaction considerations – the chemical reaction that is a reaction of a reactant exchange for the corresponding ion of another reaction. skeleton structure – two-dimensional graphic representation of atoms and links to a molecule using inserted symbols and solid lines for links. Sodium - Sodium is the name for the element with atomic number 11 and is represented by the symbol Na.sol - the sloom type of that solid particle to stop in a liquid. solid - state of subjects characterized by high degrees of organization, and a stable and volume form. solidification – the phase of change that results in the formation of a solid. Solcibility – maximum number of salvation that can be dissolved into a specified solution. Product solubilities – KSP, the constant poised for a chemical reaction in which a solid compound ionic dissolves its yield ions into solution. solid - dissolved substance in a chemical solution. solution – the homogeneous mixture of two or more substances. solvent - component of a solution present in the greater proportion. specific gravity – ratio of the density of a substance in the water of water. specific heating – the amount of heat required to increase the temperature of a mass of a specified amount. specific heat capacity – the amount of heat needed to increase the temperature of a substance per mass unit. ion - ion spectiators were found in the same amount on both the reactive and generated location of a chemical reaction that does not affect poised. spectroscopy – analysis of the interaction between problems and any portion of the electromagnetic spectrum. spectrum - the characteristic length of electromagnetic radiation emitted or absorbed by an object or substance. number proportions turn (ms) - fourth quantity, which indicates orientation at times intrinsic moment angular to an electron in an atom. Spontaneous fission - divided spontaneously into an atomic nuclear at two smaller nuclei and usually neutrons, accompanied by the release of energy. Spontaneous processes – processes that can happen without any energy input from the environments. standard – reference used in calibrated measurement. standard hydrogen electrode – SHE, the standard measurement of electrodic potential for the thermodynamic scale of redox potential. standard oxidation potential – the volt potential generated by a half-reaction oxidation compared to standard hydrogen hydrogen at 25°C, 1 in pressure with a concentration of 1 M. standard potential reduction – the potential of volts produced by a half-reaction reduction compared to electrogen hydrogen standards at 25°C, °C 1 in pressure and a concentration of 1 M. standard solution – a solution with a precisely known concentration. standard temperature and pressure – STP, 273 K (0° Celsius or 32° Fahrenheit) and 1 at pressure. state of question – homogeneous phase of matter (e.g., solid, liquid). Steam distillation – distillation process in which steam or water is added to lower boiling points of compounds. Steel - an alloy iron with carbon. sterical number – the number of atoms linked to a central atom of a molecule plus the number of electronic pairs attached to the central atom. stock solution – focus solution intended for divided at a lower concentration for actual use. stoichiometry - study of quantitative relationship between substances that underwent a physical or chemical change. STP – standard temperature and pressure; 273 K (0° Celsius or 32° Fahrenheit) and 1 at pressure. solid strong acid - the fully dissociated acid in its ions in aqeluze solution. strong base – base that is fully dissociated in its ions in aqent solutions (e.g., NaOH). strong electrolytes - electrolytes that are completely dissociated in aqent solutions. Strontium – Earth’s alcohol with component symbol Sr with 38 atomic numbers. sublimation – transition phase from solid phase directly to steam phase. subshell - subdivision of electron shell separated by electron orbitals (e.g., s, p, d, f). subtraction – the medium on which a reaction occurs or reactive offers a surface for absorption. substitution - atom or functional group that replaces a hydrogen atom in a hydrocarbon. Reaction to substitution – chemical reaction in which a functional group or atom is replaced by another functional group or atom. Sulfur - Sulfur is the name for the element with atomic number 16 and is represented by the S. supernate symbol – the liquid result of a precipitation reaction. supersaturated – supercooled; The condition that was a liquid cooled to a temperature below that crystallization normally occurs, yet without solid training. Surface tension – physical properties equal to strength per unit needed to expand the surface of a liquid. surfactant – species that act as a wet agent to lower liquid surface tension and increase spreads. Suspension - the heterogeneous mixture of solid particles in a liquid. reaction synthesis - direct combination reaction; chemical reactions in which two or more species combine to form a more complex product. Titan is a useful transition metal. Krischan D. Rudolph/Getty Images Tantal - metal transitions and inserted symbol Ta with atomic number 73. techniques - solid transition symbol element Tc with atomic number 43. Diolurium – metalloid and symbols ground elements with 52 atomic numbers. temperature - property of problems which is a kinetic energy measurement of its particles; heat or cold. Terrible - ground elements with TB symbols and 65 atomic numbers. tetrahedral - molecular geometry in which a central atom forms four links direct toward the corners of a regular tetrahedron. Maverick carbon - a carbon atom forming five covalent convalents, forming a structure that resembles a star. thallium – metal with 81 atomic numbers and component symbols Tl. theoretical yield – the number of products that should be found if the limit is limited to a fully reacted reaction reaction. Theory – a well-established explanation of scientific data that can be disproved by a single contrary result. thermodynamics – scientific studies of heat, labor, and related properties of mechanical and chemical systems. plastic thermoset - a polymer that was made endurancefully rigid about heating. thiol - an organic sulfur compound that includes an alkyl or anaryl group and a sulfur-hydrogen band; R-SH. thiol group - functional group containing a sulfur tied to a hydrogen, -SH. thorium - Thorium is the name for the element with 90 atomic number and is represented by the th symbol. thulium - rare earth element with 69 atomic number and component symbol Tm. tin - metal with 50 atomic numbers and Sn. tinctoring element symbols - an extract of a sample into a solution, usually with alcohol as the solvent. Titanium - Transition metal and small element symbol with 22 atomic numbers. title - solution of known concentrations used in a attraction to determine the concentration of a second solution. titleration – process for adding a known volume and concentration to one solution to another determines the concentration of the second solution. Torr – unit of pressure is equal to 1 mm Hg or 1/760 atmospheric pressure standard. transomer - the isomer in which functional group occurs on opposite sides of the double bond. transition interval - concentration range of chemical species that can be detected using an indicator. metal transition - components from group B to the periodic chart characterized are not partially full d electronic sublevels orbital. translational energy – energy in the space movement. transmit - switch from one shape or substance to another. triple point – temperature and pressure of that solid, liquid, and vapor phase of a substance coexist to stabilize with each other. Metal transitions with 74 atomic numbers and symbols of the element W. Tyndall effect - the spread of a beam of light as is passed to a colloid. Ultraviolet light is sometimes called dark light because it's beyond the visible spectrum. Cultura RM Exclusive / Matt Lincoln / Getty Images ultravoleit radiation - increase electromagretic radiation a padn between 100 nm and 400 nm. Sometimes they call out dark light. malfunction ID – a four-digit card code used to identify dangerous or flammable chemicals. U.N. identifier number – a U.N. ID used to transport dangerous materials. unit - a standard used for comparison of measurements. universal constant fuel - usually indicated by R, the constant of gas is the Boltzmann constant in energy unit per molester: R = 8.3145 J/mol · Which universal indicator - a mixture of pH indicators used to measure pH more than a wide range of values. universal solvent - a chemical product that melts more substances. While water is often called the universal solvent, most nonpolar molecules are insoluble to it. unusual fat - a lipid with no dual carbon-carbon link. unsaturated solution – a solution through which solc concentration is lower than its solcibility. All solutions featuring dissolved in the solution. uranium - 92 components and symbols U.Volumetric Flask are used to prepare chemical solutions. COLIN CUTHBERT/SCIENCE PHOTO LIBRARY/Getty Images vacuum – a volume with little to no problem (no pressure). valence - the amount of electrons needed to fill the electron valve outward. valence bond theory – explanation of bonding between two atoms as a result of the overlap of half-full orbital atomic. Electron valence - electrons that are most likely to participate in bondage training or a chemical reaction. Valence Shell Electron Pe Repulsion Theory - molecular model that predicts the atom geometry of a molecule by minimizing electrostatic forces between electrons valen around a central atom. vanadium - Vanadium is the name for the element with atomic number 23 and is represented by the symbol V. It is a member of the Transition Metal Group. Van der Waals strength - weak force that contributes to intellectual bonds. Van der Waals rays - half the distance between two hanging atoms in a state of electrostatic balance. vapor - a condensable gas. vaporated pressure – pressure exercised by a vapor of poised with liquid or solid phase of the same substance or the partial pressure of a vapor above liquid or solid. vaporization - transition phase from the liquid phase to gas phase. vector - a geometric object that has both greattness and direction. viscosity – how liquids a liquid flow, which is the ratio between a tailor-applied stress and gradient at the resulting speed. visible light - electromagiere radiation that can be known by the human eye, usually from 380 nm to 750 nm (400 to 700 nm). volatile – a substance that is easily vaporized. volume - the three-dimensional space occupied by a solid, liquid, or gas. baggy volumes - the type of glass chemistry used prepare known concentration solutions. volume-volume percentage – v/v% is the ratio between the volume of a substance in a volume solution of the total volume of the solution, multiplying by 100%. VSEPR – see Valence Shell Electron Pe Repulsion Theory called the solvent universal because many dissolved compounds in it. Yuji Sakai/Getty Images Water - a compound formed by an oxygen atom and two hydrogen atoms. This usually refers to the liquid shape of the molecule. Water gas - a combination gas of hydrogen gas and carbon monoxide. water of crystallization - the water is stoichiometrically tied to a crystal. Water in hydration - stoichiometrical water is tied to a compound, forming a hydrated. waves – a function that describes the probability of the amount state of a particle in terms of turn, time, position, and/or momentum. wave - the distance between identical points in two successive waves. Wave-particle duality – concept that photon with submortal particles display properties in both waves and particles. sir - a lipid that includes ranges of esters or alkanan from fat acids and alcohol. weak aside - an aside which only partially dissociates in its ions of water. weak base - a basis that only partially disruption of water. weak electrolytes - an electrolyte that does not completely dissociate in its ions of water. edge-and-dash projection – molecule representation using three line types to show three-dimensional structures. weight – the strength on a mask due to the acceleration of gravity (mass multiplied by acceleration). word equation – a chemical equation expressed in words rather than chemical formulas. labor - force multiply by distance or the amount of energy needed to move a mask against a force. working solution – a chemical solution prepared for use in a lab, usually by diluting a stock solution. Xenon is often found in plasma balls. DAVID PARKER / SCIENCE PHOTO LIBRARY / Getty Images xenon - Xenon is an element with an atomic number of 54 and atomic weight of 131.29. It is a gas odorless energy used to fill cathode ray tubes. X-ray radius are light rays with a length from 0.01 nanometers1.0. Also known as: X Radiation Ytrium is one of the rare earth elements. DAVID MACK / Getty Images Yield – In chemistry, yield refers to the amount of a product obtained through a chemical reaction. Chemistry refers to experimental yield, actual yield, theoretical yield rates, and the percent yield differentiated between calculating yield values and people who actually get out of a reaction. ytterbium – Ytterbium is element number 70 with a Yb element symbol. ytrium - Yttrium is an element element with an atomic number of 88.90585. It’s a dark gray metal that is used to make alloys for nuclear technology because the component has a high transparency. Zinc is one of transition metals. Bar?s Muratoglu/Getty Images Zaitsev rule - rule of organic chemistry that states training alkene from an elimination reaction will produce most highly replaced alcans. zeta potential (ζ-potential) – the potential difference across the boundary phase between a liquid and a solid. Zinc - Zinc is the name for the element with atomic number 30 and is represented by the symbol Zn. A transition metal group member. Zirconium - Zirconium is the name for the element with 40 atomic number and is represented by the symbol Zr. Is a member of the transition metal group. Group.

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