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## Modeling chemistry unit 5 review answers

Unit 5 - Counting particles too small to see instructional targets 1. Use the Avogadro hypothesis and experimental data to determine the relative mass of molecules. 2. Use experimental data to determine the number of items in the sample without actually counting them. 3. Use experimental data to determine the molar mass of the elements. 4. Taking into account the chemical formula of the substance, the molar mass should be determined. 5. Given the mass of the substance, determine the number of sample moles. number of atoms or molecules in the sample. 6. Given the number of moles of the substance, find the weight of the sample. number of atoms or molecules in the sample. 7. Taking into account the model of the compound, determine its % composition. 8. Taking into account the data on % of the composition of the sample, an empirical formula of the compound should be defined. 9. Taking into account the empirical formula and information on the molar mass of the compound, a molecular formula should be determined. Unit 5 Tutorials Conversions: Mass - Moles - Particles Empirical and Molecular Formulas: What Are They? Empirical and molecular formulas: Finding one of the other. Unit 5 Documents Unit 5 WS1: Molar Mass Elements Unit 5 Notes Sheet Unit 5 Student Reading for Notes Unit 5: Size of Mole WS Unit 5 WS 2: Molar Mass, Moles, and Molecules Unit 5 WS 3: Empirical and Molecular Formulas Unit 5 Review WS Thank you for the most interesting on our sites. We are a non-profit group that runs this website to share documents. We need your help in maintaining this website. For our site to work, we need your help to cover the costs of our server (about 400 USD/m), a small donation will help us a lot. Help us share our service with your friends. [Loading Livebinder](#)

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