

CONCEPT REVIEW  
**9**

# CHEMICAL CHANGES

PAGES 50 TO 58

Complete this concept review handout and keep it as a record of what you have learned.

## DEFINITION

A chemical change \_\_\_\_\_  
\_\_\_\_\_  
\_\_\_\_\_

## CHEMICAL REACTION OR CHEMICAL CHANGE

- The law of conservation of mass states: \_\_\_\_\_  
\_\_\_\_\_  
\_\_\_\_\_
- During a chemical reaction one or more substances, called \_\_\_\_\_, react to make new substances called \_\_\_\_\_.
- During a chemical change, the number of atoms stays \_\_\_\_\_

## TIPS TO HELP RECOGNIZE CHEMICAL CHANGES

- \_\_\_\_\_  
\_\_\_\_\_
- \_\_\_\_\_  
\_\_\_\_\_
- \_\_\_\_\_  
\_\_\_\_\_
- \_\_\_\_\_  
\_\_\_\_\_
- \_\_\_\_\_  
\_\_\_\_\_



EXAMPLE OF A CHEMICAL REACTION

$H_2 + Cl_2 \rightarrow 2 HCl$		
Before the chemical reaction		After the chemical reaction
Reagents	Number of atoms	Product Number of atoms
$H_2 + Cl_2$	2 atoms of H	
	2 atoms of Cl	

TYPES OF CHEMICAL CHANGES

Chemical change	Definition	Energy use
_____	_____ _____ _____	_____ _____ _____
_____	_____ _____ _____	_____ _____ _____
_____	_____ _____ _____	_____ _____ _____
_____	_____ _____ _____	_____ _____ _____