

Counting Aluminum Atoms- An investigation into the size of atoms

Your group has an aluminum cube that has a mass of 44 grams. A chemistry book told you that the density of aluminum is 2.6 g/cm^3 . Another said that for every 27 grams of aluminum there are 6.02×10^{23} atoms. Use this information to answer the following questions.

1. What is the volume of the aluminum cube?
2. How many aluminum atoms are in the cube?

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2. How many aluminum atoms are in the cube?

Counting Brass Atoms- An investigation into the size of atoms

Your group has a brass cube that has a mass of 139 grams. A chemistry book told you that the density of brass is 8.6 g/cm^3 . Another said that for every 129 grams of brass there are 6.02×10^{23} atoms. Use this information to answer the following questions.

1. What is the volume of the brass cube?
2. How many brass atoms are in the cube?

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Your group has a brass cube that has a mass of 139 grams. A chemistry book told you that the density of brass is 8.6 g/cm^3 . Another said that for every 129 grams of brass there are 6.02×10^{23} atoms. Use this information to answer the following questions.

1. What is the volume of the brass cube?
2. How many brass atoms are in the cube?

Counting Iron Atoms- An investigation into the size of atoms

Your group has an iron cube that has a mass of 126 grams. A chemistry book told you that the density of iron is 7.8 g/cm^3 . Another said that for every 56 grams of iron there are 6.02×10^{23} atoms. Use this information to answer the following questions.

1. What is the volume of the iron cube?
2. How many iron atoms are in the cube?

Counting Iron Atoms- An investigation into the size of atoms

Your group has an iron cube that has a mass of 126 grams. A chemistry book told you that the density of iron is 7.8 g/cm^3 . Another said that for every 56 grams of iron there are 6.02×10^{23} atoms. Use this information to answer the following questions.

1. What is the volume of the iron cube?
2. How many iron atoms are in the cube?