1. Set up conversion factor \( \frac{\text{mole A}}{\text{mole B}} \).
2. Multiply \( A \) by \( \text{mole of B} \).
3. Solve for \( \text{mole of unknown} \).
4. Balance the equation.
Using the masses of elements:

\[
\text{mass of element} = \text{mole of element} \times \text{molar mass}
\]

Use the molar ratio of the balanced equation:

\[
\text{mol A : mol B} = \frac{\text{Lites of A}}{\text{Lites of B}} \times \frac{22.4\text{L}}{1\text{mol}}
\]

Lites of A: 1L

Mols of B

Mols of A